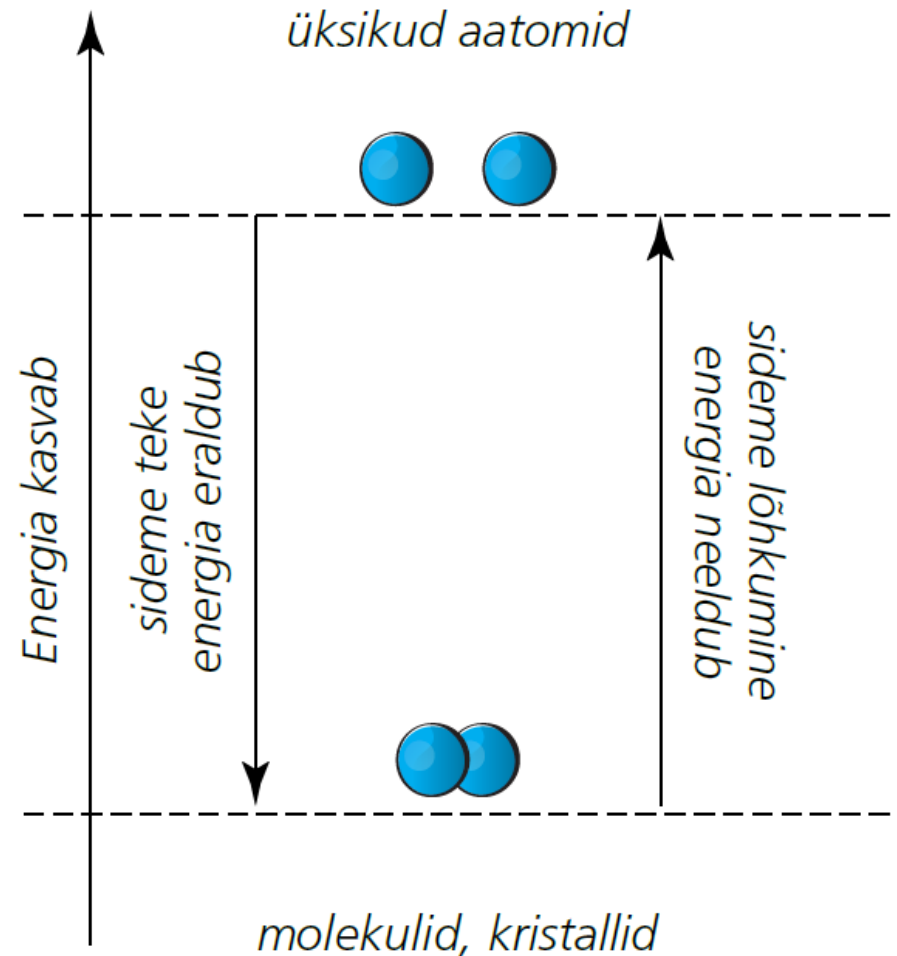


Kovalentne side

Neeme Katt

Energia miinimumi printsiip

- Aatomite liitumisel molekuliks lähevad nad püsivamasse (madalama energiaga) olekusse
 - Elektronidega täielikult täitunud väliskiht (taval $8e$)
- Keemiline side on mõju, mis hoiab aatomeid molekulis koos

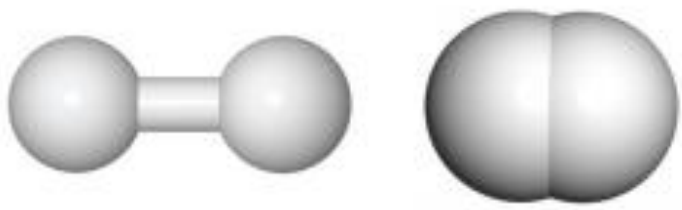
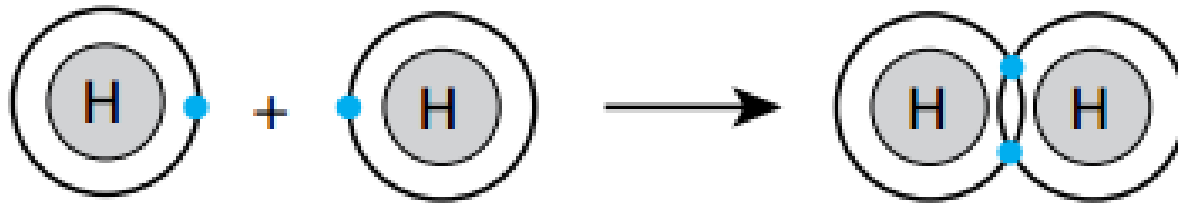


Kovalentne side

- Molekulid koosnevad aatomitest, mis on omavahel seotud kovalentse sidemega
- **Kovalentne side** on ühiste elektronipaaride abil tekkinud side
 - Ühine elektronipaar tekib väliskihi üksikutest elektronidest
 - Esineb mittemetalli aatomite vahel

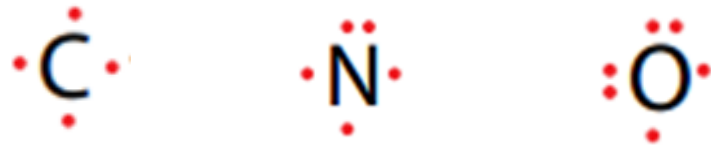
<http://mudelid.5dvision.ee/keemside>

Vesinik H₂

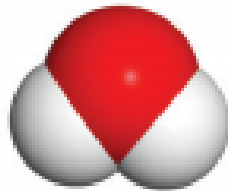
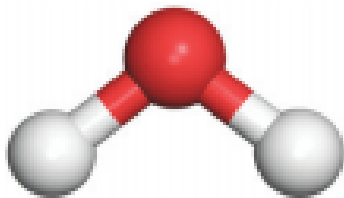
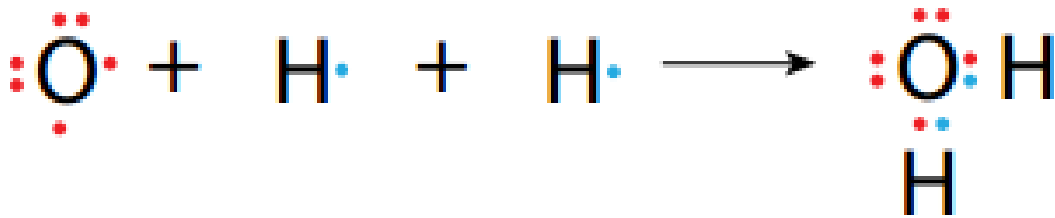
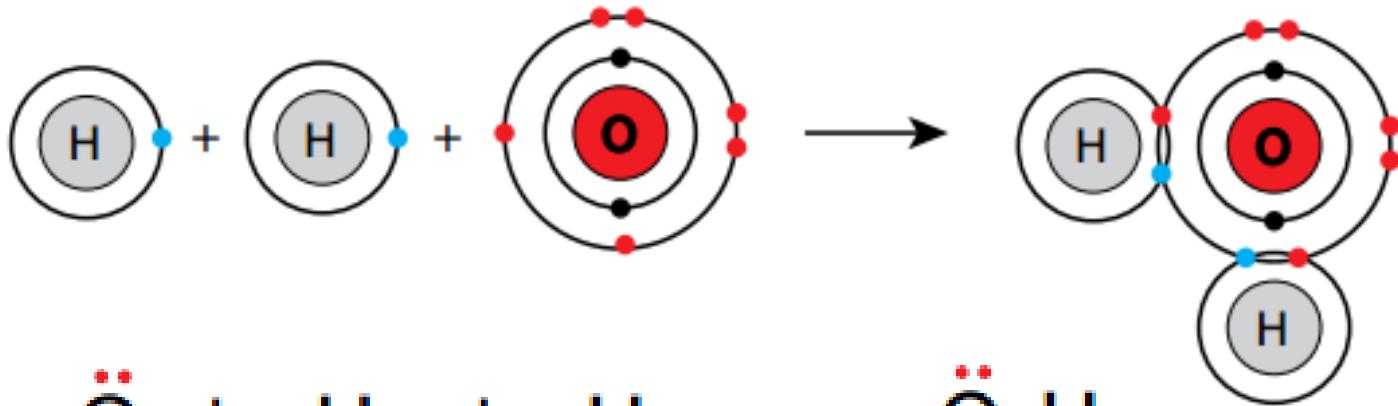


Kui palju üksikuid elektrone on aatomi väliskihil?

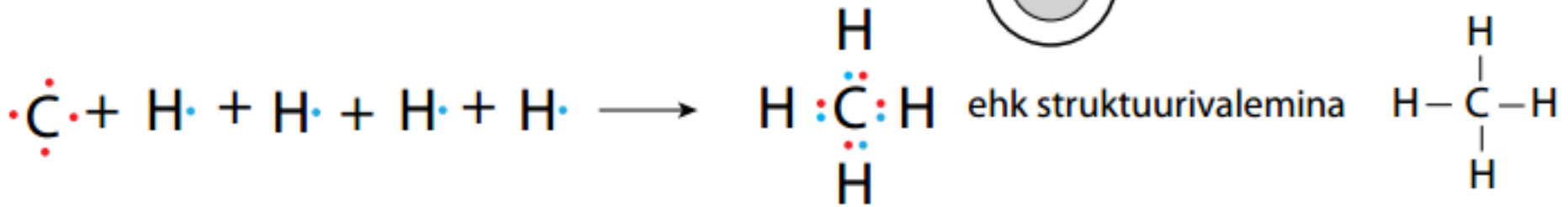
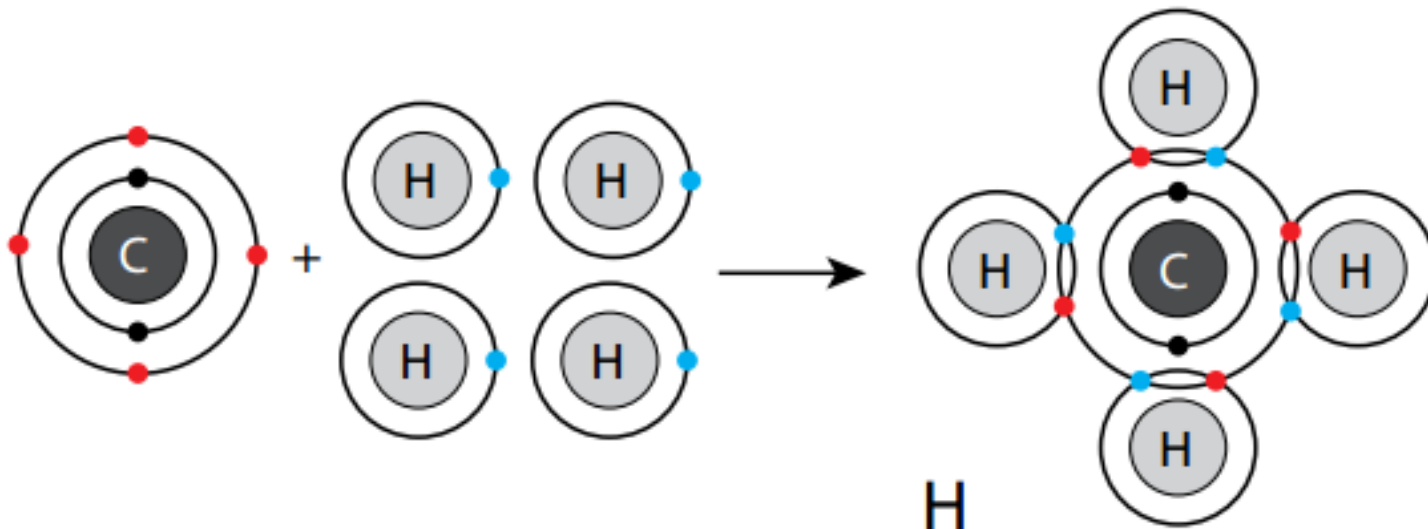
- Väliskihil võib olla kuni 8 elektroni
- Väliskihi elektronid võivad moodustada paare või olla üksikult.
 - kui on vabu toole, siis istutatakse üksi;
 - kui vabu kohti enam pole, siis tuleb kellegagi tooli jagada (moodustada paar).



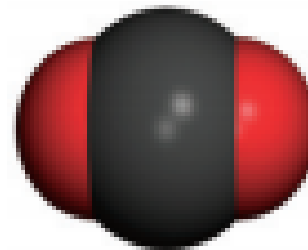
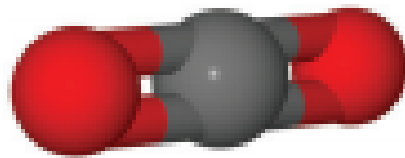
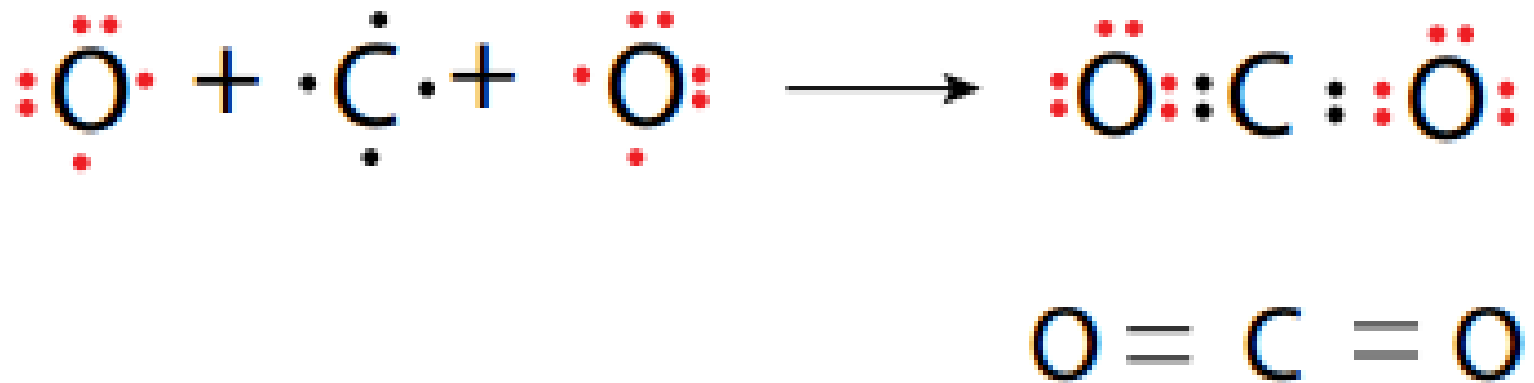
Vesi H₂O



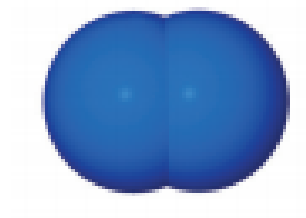
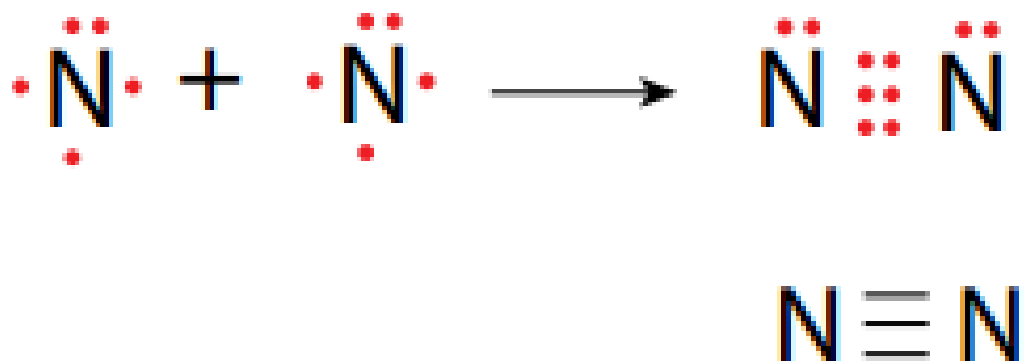
Metaan CH₄



Süsinikdioksiid CO₂



Lämmastik N₂



Paratsetamool $C_8H_9NO_2$

